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mass in grams of 15×10^{16} atoms s.
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reacts with oxygen to produce
aluminum oxide as follows: $4\text{Al} + 3\text{O}_2$

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? $2\text{Al} + 3\text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ a. If you use 2.3 moles of Al, how many moles of Al_2O_3 can you make? b. If you want 3.9 moles of Al_2O_3 , how many moles of O_2 are needed?

2. In the presence of sulfuric acid, metallic iron forms iron(III) sulfate: $2\text{Fe} + 3\text{H}_2\text{SO}_4 \rightarrow \text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2$

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chemistry dealing with the numerical
relationships in chemical reactions .

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