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Solutions: homogeneous mixtures of two or more pure substances, having uniform composition and properties throughout.

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Electrochemistry It is the study of production of electricity

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from I energy which is released during spontaneous chemical reactions, also the use of electrical energy to bring about non-spontaneous chemical transformations.

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Electrons released at anodic spot move through metal and go to another spot on the metal and reduce  $\text{O}_2$  in presence of  $\text{H}^+$ . This spot behaves as cathode.  $\text{O}_2(\text{g}) + 4 \text{H}^+(\text{aq}) + 4 \text{e}^{-} \rightarrow 2 \text{H}_2\text{O(l)}$   $E^{\circ} = 1.23 \text{ V}$ . The overall reaction is,  $2 \text{Fe(s)} + \text{O}_2(\text{g}) + 4 \text{H}^+(\text{aq}) \rightarrow 2 \text{Fe}^{2+} + 2 \text{H}_2\text{O(l)}$   $E^{\circ} = 1.67 \text{ V}$

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Solution A homogeneous mixture of two or more substances is known as a solution. Solute : The substance present in smaller amount in a solution is called solute. Solvent: The substance present in larger amount in a solution is called solvent.

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Kerala Plus Two Chemistry Notes Chapter 9 Coordination Compounds. Coordination chemistry – branch of chemistry which deal with the complex compounds formed by transition and other metals. Chlorophyll, haemoglobin and vitamin B 12 are coordination compounds of Mg, Fe and Co respectively. Werner ' s Theory of Coordination Compounds The main postulates are,

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$2\text{Ca}(\text{OH})_2 + 2\text{Cl}_2 \rightarrow \text{Ca}(\text{OCl})_2 + \text{CaCl}_2 + 2\text{H}_2\text{O}$   $\text{CaCl}_2$  is a

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powerful bleaching agent.  $\text{Cl}_2 + \text{H}_2\text{O} \rightarrow 2\text{HCl} + [\text{O}]$

Coloured substance +  $[\text{O}] \rightarrow$  colourless substance. Uses: For bleaching wood pulp, cotton and textiles; for the preparation of insecticide, pesticides and other organic solvents, e.g.  $\text{CHCl}_3$ , DDT, BHC etc. Hydrogen Chloride (HCl):  
Preparation:

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