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against (temperature in Kelvin) for the second order decomposition of into and O. a.State how the rate constant, varies with temperature, . [1]  
b.Determine the activation energy, , for this reaction. [3] The rate expression for this reaction is rate and the rate constant is at 750 °C.

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°C with hydrogen to form nitrogen and water. All reactants and products are in the gaseous phase. The gas-phase decomposition of dinitrogen monoxide is considered to occur in two steps. The experimental rate

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expression for this reaction is rate . The conversion of into is an exothermic reaction which can be represented as follows.

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