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Ans. Acetic acid dissociates to produce one  $H^+(aq)$  ion per molecule and hence its basicity is one.  $CH_3COOH + H_2O \rightleftharpoons H_3O^+ + CH_3COO^-$   
Magnesium hydroxide reacts completely with  $2H^+(aq)$  ions of an acid to form salt and water, therefore its acidity is 2.  $Mg(OH)_2 + H_2SO_4 \rightarrow MgSO_4 + 2H_2O$   
11. Why is sulphuric acid a dibasic acid? Give three reasons.

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from the molecule of an acid, with metal ions, is called an acid salt. For example, 1. Sodium hydrogen sulphate ( $NaHSO_4$ ) 2. Calcium hydrogen carbonate [ $Ca(HCO_3)_2$ ]. (iii) Basic salt - A salt formed by the partial neutralisation of hydroxyl ions ( $OH^-$ ) of a base, by an acid is called a basic salt. For example,

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Download Free Acids Bases Salts Testlabz hydrogen atoms are called tribasic acids. 2. The substances which react with acids to form salt

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and water as the 4 ACID, BASES AND SALTS - Testlabz Acids, Bases and Salts 154 SCIENCE AND TECHNOLOGY Notes MODULE - 2 Matter in our Surroundings 8 ACIDS, BASES AND SALTS From generations, our parents

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Acids, bases and salts - (CCEA) Many chemicals are acidic, neutral or alkaline. We can distinguish one from another using indicators. Acidity and alkalinity are measured on the pH scale.

~~Acids, bases and salts - Acids, bases and salts - (CCEA ...~~

18 Science 2.12.12.1 UNDERSTANDING THE CHEMICAL PROPERTIES OF ACIDS AND BASES 2.1.1 Acids and Bases in the Laboratory Activity 2.1 These indicators tell us whether a substance is acidic or basic by change in colour. There are some substances whose odour

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changes in acidic or basic media. These are called olfactory indicators.

## ~~CHAPTER2 Acids, Bases and Salts~~

Acids, Bases, and Salts 40,236 Many acids and bases occur naturally in nature, such as citric acid in fruits like orange, lemon etc, tartaric acid in tamarind, malic acid in apples and lactic acid in milk and milk products, hydrochloric acid in gastric juices. Similarly, many bases are found such as lime water.

## ~~Acids, Bases, and Salts — Introduction, Dissociation ...~~

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Some examples of bases are: (i) Sodium hydroxide (NaOH) or caustic soda used in washing soaps. (ii) Potassium hydroxide (KOH) or potash

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used in bathing soaps. (iii) Calcium hydroxide ( $\text{Ca}(\text{OH})_2$ ) or lime water used in white wash. (iv) Magnesium hydroxide ( $\text{Mg}(\text{OH})_2$ ) or milk of magnesia used to control acidity.

### ~~Notes ACIDS, BASES AND SALTS~~

Acids, bases and salts - Higher. Acid solutions can be dilute or concentrated. Concentration is calculated from mass of solute and volume of solvent. Acids may be weak or strong depending on how ...

### ~~Acids, bases and salts - Higher - Edugas test questions ...~~

The salts of strong acids and weak bases give acidic solution having pH less than 7. Example,  $\text{NH}_4\text{Cl}$ , Ammonium Chloride will have pH less than 7. The salts of weak acids and strong bases give basic solution having pH more than 7. Example,  $\text{Na}_2\text{CO}_3$ , Sodium Carbonate will have pH more than 7. Question 7.

### ~~Acids, Bases and Salts Class 10 Important Questions with ...~~

Arrhenius Definition of Acids and Bases An acid is a substance that reacts with water to produce hydronium ions,  $\text{H}_3\text{O}^+$ . A base is a substance that produces hydroxide ions,  $\text{OH}^-$ , in water.

CHEM 0012 Lecture Notes 3

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## ~~Acids, Bases and Salts~~

Salts . When acid and base neutralize, salts are formed. Strong acid and strong base combines to form neutral salt.  $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$ . Eq.1. Formation of Neutral Salt. Strong acid and weak base combine to form acidic salt. For Example, Hydrochloric Acid and ammonium hydroxide combine to form ammonium chloride. Other examples, sodium hydrogen carbonate, sodium hydrogen sulphate etc.

## ~~Revision Notes on Acids, Bases and Salts — Askiitians~~

Acid + base  $\rightarrow$  salt + water + heat.  $\text{H}_2\text{SO}_4 + \text{MgO} \rightarrow \text{MgSO}_4 + \text{H}_2\text{O}$   $2\text{HCl} + \text{Mg}(\text{OH})_2 \rightarrow \text{MgCl}_2 + 2\text{H}_2\text{O}$ . 2. Reaction of non-metal oxides with bases. Non-metal oxides are acidic in nature Base + Non-metal oxide  $\rightarrow$  salt + water + heat.  $2\text{NaOH} + \text{CO}_2 \rightarrow \text{Na}_2\text{CO}_3 + \text{H}_2\text{O}$ . To know more about Properties of Acids and Bases, visit here. Water Acids and bases in water

## ~~Class 10 Chapter 2 Science Notes Acids, Bases, and Salts~~

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